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**Selective and quantitative nitrate electroreduction to ammonium using a porous copper
electrode in an electrochemical flow cell**

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Abstract

The aim of this work was to set up a novel electrochemical system allowing an efficient transformation of concentrated nitrate solutions to ammonium and which can be subsequently implemented on a large scale application. First, this paper describes the preparation of a porous copper modified electrode by successive electrodeposition of nickel then copper on a graphite felt of large specific surface area. Homogeneous Cu coating of all fibres in the 3D porous structure was successfully obtained using low concentrations of copper salts and high applied current intensities. The porous copper electrode was then used in a flow electrochemical process to achieve a selective and quantitative transformation of concentrated nitrate into ammonium. Different electrolytic solutions, slightly acid (acetate buffer) or neutral (phosphate buffer), and flow rates were investigated. The nitrate solution was quantitatively reduced into NH_4^+ with high selectivity in only one pass through the electrode. When the applied current was similar to the theoretical one, the maximum selectivity (96 %) and the best current efficiency (72%) for NH_4^+ formation were reached at pH 7.2 with a flow rate of 2 mL min^{-1} . The obtained ammonium solution can be subsequently used either as a potential nitrogen source during microbial culture or simply as a fertilizer.

Keywords: porous copper cathode; nitrates electroreduction; ammonium; electrochemical flow cell

1. Introduction

Nitrate ions (NO_3^-) contamination of ground and surface water resources has become a serious problem around the world. Removing nitrate is very important for various reasons. (i) Nitrate water pollution disturbs public health and ecological equilibrium [1]. Consuming high NO_3^- concentrations in drinking water can cause serious health problems in humans such as the “blue baby syndrome” in infants, liver damage and cancer [2]. (ii) By considering the ecological balance, excess of nitrate favor eutrophication of lakes, rivers and seas, manifested through uncontrolled algal growth [3,4]. WHO and European community set stringent limits for the NO_3^- concentrations in surface and groundwater to 50 mg L^{-1} and 25 mg L^{-1} respectively [5].

Nowadays, many processes such as biological and physiochemical methods are used for nitrate removal from contaminated water but show drastic limits. Bionitrification is considered as the method of choice [6,7]. However, this method is slow, requires continuous monitoring of microorganism's growth (carbon source, pH and temperature control) [8]. Furthermore, it can be used only at concentrations lower than $1000 \text{ mg NO}_3^- \text{ L}^{-1}$, since higher ones can be poisonous to the bacteria [9]. Physicochemical processes such as ion exchange, reverse osmosis and electrodialysis are costly [10] and produce secondary brine wastes, because the nitrate are only separated but not destroyed [11]. Thus, these methods lead to concentrated effluents of nitrates ($\geq 3 \text{ g L}^{-1}$) [12] that are very difficult to treat.

Electrochemical nitrate treatment can offer a promising and attractive solution due to its advantages regarding low cost effectiveness, convenience, ability to treat highly concentrated nitrate effluents and environmental friendliness [13–15] particularly if the resulting product is N_2 gas [16]. An ideal electrochemical method must combine a high rate of reduction, a selective conversion of nitrate to the desired product, a high current efficiency (% CE) and low energy consumption [17].

Almost all recent electrochemical studies aim to convert nitrate to the non toxic nitrogen gas [18]. In this context paired electrolysis seems to be the most efficient and hence the most widely used electrochemical method [6,19–21]. It consists to reduce NO_3^- to ammonium at the cathode, and then the produced ammonium is oxidized at the anode into N_2 with or without the contribution of hypochlorite ions (ClO^-), which results from the oxidation of chloride ions on the anode [22]. However, the main problem in this approach is that the transformation of NO_3^- to nitrogen is a difficult reaction since both the rate and the selectivity of the reduction of nitrate to ammonium are usually low, due to the large number of by-products such as NO_2 , NO_2^- , NO , N_2O , NH_2OH , NH_3 , NH_2NH_2 involved in the reduction mechanism [18].

As the products of the nitrate electroreduction are critically dependent on the nature of the cathode material [23–25], this fact led to the search for selective and active cathode materials. Many solid electrodes have been studied including Pb, Ni [26], Cu [27], Pt [28], Ru [29], Ir [30], Rh [31] and Fe [32]. Among these materials, copper is known to be the most efficient electrocatalyst for nitrate electroreduction producing ammonium as a final product [33–35]. Thus, electrochemical treatment of nitrates on pure Cu cathodes [33,36,37] or Cu alloys [21,38–40] has been widely studied. But all these investigations are analytical studies and/or deal with relatively low nitrate concentrations [39]. Moreover, the main problem in the electroreduction of nitrates on solid electrodes is the low reduction rate [7] owing to the low mass transport. This considerable drawback limits their use for large scale applications.

The use of a flow electrochemical cell with a graphite felt as porous electrode can resolve the problem related to the mass transport. This system has the advantage of oxidizing/reducing a large amount of substrate in short times, allowing quantitative electrochemical treatments [41]. The major advantages of graphite felt as porous electrode material are its high specific surface area, ranging from $0.7 \text{ m}^2 \text{ g}^{-1}$ to $2000 \text{ m}^2 \text{ g}^{-1}$, the low cost and the availability of these materials.

But unfortunately, graphite felt is a poor electrode for the selective reduction of nitrates to ammonium. Hence, it's necessary to enhance its selectivity toward ammonium formation.

It is interesting to note that among the various nitrate electroreduction by-products (ie. N_2 , N_2O , N_2O^- , NH_4^+), ammonium has the highest economic value. It has an important potential as fertilizers [42–44] and as nitrogen source for microorganism cultures producing biohydrogen [45–47]. The use of cheap inorganic nitrogen source appears relevant from an economical point of view, contrarily to organic nitrogen sources which are often more expensive [48]. For this reason, an electroreduction of nitrate to NH_4^+ over copper cathode can be considered as an attractive way, since it allows not only pollutant treatment, nitrate, but also possible valorization of the treated effluent.

Therefore, the first objective of the present study was the preparation of a new cathode material, a porous copper electrode by successive electrodeposition of nickel then copper on a graphite felt. A homogeneous coating of all fibres of thick graphite felts was obtained. The method leads to the formation of a copper modified graphite felt (CuGF) of high specific surface area and low weight. The second objective was the use of the CuGF electrode for the electrochemical reduction of a concentrated nitrate solution ($3 \text{ g NO}_3^- \text{ L}^{-1}$) into ammonium in a divided flow cell. Results show that the system combining the performances of the flow cell and the CuGF electrode lead to high chemical yields and current efficiencies.

2. Experimental procedure

2.1. CuGF working electrode: preparation and characterization

Nickel (II) sulfate heptahydrate ($Ni(SO_4) \cdot 7H_2O$); copper (II) sulfate pentahydrate ($Cu(SO_4) \cdot 5H_2O$) 99% ; boric acid 99% and sodium sulfate anhydrous 99% were supplied by ACROS. Graphite felt (RVG 4000) used as metal electroplating support was purchased from

MERSEN (Paris La Defense, France). Its specific surface area measured by the BET method was $0.7 \text{ m}^2 \text{ g}^{-1}$, its density was 0.088 g cm^{-3} and its carbon yield was 99.9%.

Graphite felt modification consisted in two steps. First, nickel was electrochemically deposited on the graphite fibres of the graphite felt (diameter: 5.6 cm and thickness 0.3 cm) according to a previously described method [49]. Secondly, the electrodeposition of copper was performed on the nickel modified graphite felt according to the conditions described in Table 1.

Table 1

1 L of electrolytic solution (copper sulfate ($3 \times 10^{-2} \text{ M}$); sodium sulfate (0.5 M); boric acid (0.25 M)) was flowed through the porous electrode in a home-made flow cell (Fig.1). To obtain a good homogeneity of the potential distribution in the three dimensional working electrode, the Ni modified graphite felt was located between two counter-electrodes (platinum grids) [50] and the electrical contact with the working electrode was performed with a copper wire. The electroplating experiments have been realized without adding cationic exchange membranes. A periodically changing current (on-pulse: 0.1 A for 1 cm³ of felt, off-pulse: 0 A, see table 2 for the duration of the pulses) was applied between the electrodes, using a power Lab/4SP (Scope version 3.6.10, AD Instruments) coupled to an EG&G Princeton Applied Research model 362 operating under galvanic conditions. The flow rate of the electrolytic solution was maintained at 5 mL min^{-1} and the pH of the electrolyte solutions was adjusted with solutions of H_2SO_4 (0.5 M) or NaOH (5 M). The full electrodeposition time was 240 min without taking into account the relaxation time (Table 2).

Table 2

The SEM (scanning electron microscope) micrographs were obtained with a Jeol 630v1F (9 kv) microscope.

2.2. Flow electrolyses

1 L of nitrate solutions (3 g L^{-1}) were prepared by dissolving potassium nitrate (KNO_3) in acetate buffer pH 4.8 ($\text{CH}_3\text{COOH}/\text{CH}_3\text{COO}^-$ 0.5 M/0.5 M) or phosphate buffer pH 7.2 ($\text{H}_2\text{PO}_4^-/\text{HPO}_4^{2-}$ 0.25 M / 0.25 M). Electrolyses were performed at constant current in a flow cell (Fig.1) equipped with two exchange membranes (cationic and anionic) (Ionac 3470 –Lanxess SAS, Courbevoie, France)

Fig.1.

The flow rate was monitored by a Gilson minipuls 2 peristaltic pump and the current by a homemade power supply (30 V/ 3A). The cell was thoroughly rinsed with distilled water before each experiment. The applied current was calculated from the Faraday law (Eq. 1)) [41].

$$i \text{ (A)} = [\text{NO}_3^-]_{t=0} \times d \times (10^{-3}/60) \times n \times F \quad (1)$$

where $[\text{NO}_3^-]_{t=0}$ is the initial concentration of nitrates (mol L^{-1}), d is the flow rate (mL mn^{-1}), F is the Faraday's constant ($96,487 \text{ C mol}^{-1}$) and n is the number of exchanged electrons ($8e^-$) for the conversion of nitrate to ammonium.

2.3. Analytical methods

Every 30 min, 5 mL of samples were collected from the electrolytic medium and filtered through a $0.45 \mu\text{m}$ membrane before analysis. NO_3^- and NO_2^- concentrations were measured using ion chromatography (Dionex120) equipped with a Dionex IonPac AS18A Anion Exchange column

and a chemical suppressor (ASR 300-ultra 4 mm), using 8 M KOH as eluent at a flow rate of 1.04 mL min⁻¹. NH₄⁺, N₂H₄ and NH₂OH concentrations in solution were spectrophotometrically determined by an UV-vis spectrophotometer (Cary-1E, Varian) as reported elsewhere [51–53]. Nitrous oxide gas (N₂O) were obtained by headspace equilibrium extraction and analyzed by μ GC/TCD (Agilent Micro GC3000 (SRA) Instruments), pure gases were used as calibration standards [54]. The pH was measured using a 9165BN pH electrode connected to an Orion-828 pH Analyzer (Orion Research Inc., Beverly, MA, USA).

2.4. Definition of the parameters

The performance of the flow cell was evaluated from the electrochemical conversion of nitrates into ammonium in one pass through the flow cell at room temperature, using the following parameters: the normalized percentage of nitrate removal or the conversion yield (% τ) (Eq. 2), the chemical yield of reduction products (ammonium and nitrite yields) (% X_i) (Eq. 3) and the selectivity based on the ammonium yield (% $S_{NH_4^+}$) (Eq. 4), [19].

$$\% \tau = ([NO_3^-]_{t=0} - [NO_3^-]_t) \times 100 / [NO_3^-]_{t=0} \quad (2)$$

$$\% X_{NH_4^+} = ([NH_4^+]_t \times 100) / [NO_3^-]_{t=0} \quad \text{and} \quad \% X_{NO_2^-} = ([NO_2^-]_t \times 100) / [NO_3^-]_{t=0} \quad (3)$$

$$\% S_{NH_4^+} = ([NH_4^+]_t \times 100) / ([NO_3^-]_{t=0} - [NO_3^-]_t) = (\% X_{NH_4^+} \times 100) / (\% \tau) \quad (4)$$

The current-denitrification efficiency (current efficiency (% CE)), calculated based on the cathodic current efficiency is defined by equation 5:

$$\% CE = (([NH_4^+]_t \times n \times F \times V_t) / (i \times t)) \times 100 \quad (5)$$

where $[X]_{t=0}$ and $[X]_t$ are the concentrations of X in solution at initial and time t (s), V_t (L) is the electrolyzed volume at time t.

All the concentrations are in mol L⁻¹ and 1 F = 96,487 C mol⁻¹.

3. Results and discussion

3.1. Galvanostatic method

Metal electrodeposition on porous material of high surface area such as graphite felt is not easy due to the strong heterogeneity of potential distribution inside the material [55,56]. At constant current, electrical potential rapidly decreases along the depth of the material. Metallization only appears at the external face of the material with formation of a metallic crust more or less rapidly according to the metal salts concentration. To avoid this phenomenon we developed a galvanostatic method well-adapted to the 3-D electrode.

Two fundamental points are necessary to achieve homogeneous coating of the fibres (i) an electrodeposition solution containing the metal salts in low concentration (around 10^{-2} M) and (ii) a strong current (around 100 mA per cm³ of felt). Thereby diffusion-limited currents are quickly reached on the external fibres limiting the rate of electrodeposition on the external face of the graphite felt. Consequently, the excess of current allows the electrodeposition on other fibres located inside the felt. As soon as the diffusion-limited current is reached on a fibre, the electrodeposition process is accelerated on another fibre located farther from the electrode surface. Due to the decrease of the metallic cations concentration in the vicinity of the fibres, hydrogen evolution occurs and becomes more and more important. An alternating current was therefore employed with long relaxation time to stop hydrogen evolution and reduce the local decrease of metallic cations concentration inside the felt.

3.2. Copper electrodeposition

Copper electrodeposition was achieved in the flow cell (Fig.1) at a constant flow rate (d) of 5 mL min⁻¹ and with a working electrode of volume $V = 7.5$ cm³. Unfortunately, first attempts to perform electrodeposition of copper on graphite felt led to non-homogeneous coating of the

fibres. SEM images showed only small and friable particles of metallic copper, scattered along the fibres whereas copper powder predominated inside the felt. This result underlines the weak adherence between metallic copper and graphite. Since copper is known to efficiently adhere to metallic surfaces, the graphite felt was first covered by a homogeneous film of nickel (Fig. 2 a) according to a previously reported method [49].

Fig. 2.

The flow rate (d) was kept to 5 mL min^{-1} and the volume of the working electrode to 7.5 cm^3 . When a constant current is applied, the maximum time allowing the complete regeneration of the solution inside the felt is $t_{\text{max}} = V / d$. Due to the convection inside the felt resulting from percolation, the regeneration time is lower than t_{max} and depends on the initial concentration of copper. With this tendency in mind, different electrodeposition experiments were performed, varying the concentration of copper and the on-pulse and off-pulse times. The most relevant one are presented in Table 2. The best conditions are given in experiment 4 corresponding to a copper concentration of $3 \times 10^{-2} \text{ M}$ with an off-pulse time (t_{off}) of 35 s and a on-pulse time (t_{on}) of 15 s. We found that when t_{off} was around $t_{\text{max}}/3$ the renewal of the copper solution inside the felt was good. Moreover, a ratio $t_{\text{on}}/t_{\text{off}}$ less than 2.3 led to a rapid decrease of the copper concentration inside the felt. Thus, hydrogen evolution became preponderant leading to an increase of the solution pH inside the felt and precipitation of copper hydroxide. In these conditions, a poor quality of copper deposition was obtained with a low current efficiency. Homogeneous copper electroplating was achieved with a current efficiency higher than 50% for a ratio $t_{\text{off}}/t_{\text{on}}$ upper to 2.3. A decrease of t_{on} led to higher values of current efficiencies but increased the full electrolysis time.

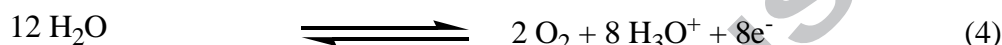
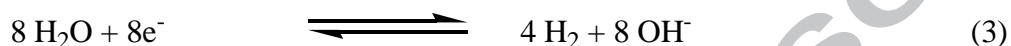
The homogeneity of the copper coating was checked by SEM. Fig. 2 b, c and d show copper coating fibres located inside the felt and cut transversally. The electrodeposition method led to a copper film with an average thickness around 2 μm that follows the shape of the graphite fibres. This layer of copper was even visible to the naked eye. As seen in Fig. 2b and c the high porosity of the material was conserved. The copper deposit has a granular aspect and consists in crystallites of different sizes assembled on the surface of the fibres (Fig. 2c). Fig. 2d shows the two successive layers of nickel and copper obtained after the two successive electrodepositions. Graphite felt is visible on the lower left quarter of the microphotograph. On the right from the bottom to the top of the microphotograph appear the nickel film on the graphite fibre and then the copper film deposited on the nickel layer. Since the brightness of nickel is very different from that of copper, SEM images clearly highlight the homogeneous structure of the copper deposit. The specific surface area of the copper modified graphite felt was estimated by BET to 2 g m^{-2} . The average density of deposited copper was about 0.265 g cm^{-3} of felt. Consequently, the electrodeposition method led to the preparation of a CuGF electrode having a high specific surface and a low density; this latter electrode was used afterward as a cathode for electrochemical denitrification studies.

3.3. Nitrate reduction

Electrolysis of nitrate was performed in a single passage in a homemade flow cell (Fig.1) containing a porous copper electrode of 5.6 cm diameter and 0.3 cm thickness corresponding to a volume of 7.8 cm^3 . The experiments were performed with a high concentration of synthetic nitrate solutions (3 $\text{g NO}_3^- \text{L}^{-1}$) to be in conditions similar to those of real effluents obtained after a classical physicochemical denitrification of polluted water [12]. The following parameters

for nitrate electroreduction were studied: (1) pH (acid and neutral); (2) flow rate (1 mL min⁻¹, 2 mL min⁻¹ and 3 mL min⁻¹); (3) applied current.

Reactions (1), (2) and (3) take place at the cathode and produce more hydroxyl anions than the protons formed at the anode (reaction 4). Consequently, the pH value increases during the electrolysis more or less strongly depending on the chemical yield of ammonium (% $X_{\text{NH}_4^+}$).



This behavior amplified by the high nitrate concentration led us to choose buffered solutions as electrolytic media. The pH of the solution was fixed to 4.8 or 7.2, since at pH higher than 8, a slight corrosion of copper occurred. Indeed, after electrolysis at pH 8, a light blue coloration of the solution appeared when the current was no more applied, due to the formation of copper salts in the solution. As soon as produced ammonia was evacuated out of the copper electrode, the coloration disappeared, showing that the corrosion process probably involved the formation of $\text{Cu}(\text{NH}_3)_4^{2+}$.

3.3.1. Effect of flow rate and pH on electroreduction efficiency

Flow rate does not seem to have a great influence on the results of electrolysis. It is obvious from table 3 that for flow rates ranging from 1 to 3 mL min⁻¹, the conversion yield of nitrate (% τ) was more than 70 % and not quantitative mainly due to hydrogen evolution (reaction 3).

Table 3

Indeed, during the electrolysis, by-products such as hydrazine ($\text{H}_2\text{N-NH}_2$), nitrous oxide (N_2O) and hydroxylamine ($\text{H}_2\text{N-OH}$) were not detected. Only the presence of a very small amount of

nitrite can be noticed (Table 3). Beyond a flow rate of 3 mL min⁻¹, the formation of gaseous hydrogen became important and prevented the electrolysis due to an accumulation of gas in the working compartment.

The selectivity of the reaction 1 was strongly influenced by the pH of the solution, as seen with % S_{NH4+} in table 3. The nitrate reduction to ammonium is quasi quantitative in neutral solution but significantly decreases in acidic solution, probably due to the formation of a small quantity of gaseous nitrogen (reaction 2) [7]. The maximum ammonium selectivity of 96 % was obtained with a flow rate of 2 mL min⁻¹. For this reason, this flow rate was selected and considered thereafter.

It is interesting to note that all experiments were performed with the same copper porous electrode and that no degradation of this material was observed, showing a high stability and durability of the electrode

3.3.2. Effect of current and pH on electroreduction efficiency

Several experiments were performed using different ratios, R_i , of applied current (i_{app}) different to the theoretical calculated intensity (i_{th}) and linked by the ratio ($R_i = i_{app} / i_{th}$). As mentioned above (section 3.3.1), all electrolyses were carried out with a flow rate of 2 mL min⁻¹ and the results are given in Table 4.

Table 4

The conversion yields of nitrate at pH 7.2 and 4.8 were similar and increased gradually with the applied current. However, it is interesting to take into account the apparent conversion yield (% τ_{app}) expressed according to the following relation:

$$\% \tau_{app} = \% \tau / ([NO_3^-]_{app} / [NO_3^-]_{t=0}) \quad (6)$$

where $[\text{NO}_3^-]_{t=0}$ is the initial nitrates concentration and $[\text{NO}_3^-]_{\text{app}}$ the theoretical maximum amount of nitrate reduced by applying i_{app} . For $i_{\text{app}} < i_{\text{th}}$ the ratio $[\text{NO}_3^-]_{\text{app}} = [\text{NO}_3^-]_{t=0}$ is equal to the ratio R_i and for $i_{\text{app}} > i_{\text{th}}$, $[\text{NO}_3^-]_{\text{app}} = [\text{NO}_3^-]_{t=0}$ and $\% \tau_{\text{app}} = \% \tau$.

The experimental results show that the lower is the applied current, the higher is the apparent conversion yield. A $\% \tau_{\text{app}}$ close to 100% was obtained when the ratio R_i was equal to 0.5, corresponding to the reduction of 50% of the initial amount of nitrate. Higher R_i values led to an increase of the amount of the nitrate reduced but decrease the $\% \tau_{\text{app}}$ (Table 4). This is due to the apparition of the electrochemical reduction of water, leading to hydrogen evolution. The behavior of the selectivity was first similar for the two solutions (pH 4.8 and 7.2) for $R_i = 0.5$ and increased from for $R_i = 0.5$ to 1. For higher values of R_i , the selectivity was remained the same for pH 7.2 but, decreased slightly for pH 4.8. Indeed, in neutral solution, the selectivity was close to 100 % for $R_i \geq 1$, showing that the secondary reaction 2 no more occurred. In acidic solution, a maximum selectivity of 75% was obtained for $R_i = 1$, due to the formation of a constant amount of gaseous nitrogen.

The current efficiency calculated from the reduction of nitrate into ammonium was maximum for R_i close to 0.9 (Fig. 3) and equal to 72 % and 50 % for pH 7.2 and pH 4.8, respectively.

Fig. 3.

As seen above, the formation of nitrogen (reaction 2) and hydrogen (reaction 3) are the main secondary reactions. In acidic solution, whatever the value of R_i , nitrogen formation always happened, decreasing the current efficiency. But in neutral solution, nitrogen formation stopped when R_i reached the value of 1 and then only hydrogen formation limited the current efficiency.

Thus, the maximum value of current efficiency for $R_i = 0.9$ corresponds to the end of nitrogen formation. Beyond this value, the formation of hydrogen is more and more important and affects the current efficiency (Fig.3).

It is important to note that the conversion yield and current efficiency are mainly limited by contact probability between nitrates ions and metallic fibres. Consequently, they depend on the porosity and the surface area of the electrode. Thus, in neutral solution, the current efficiency should be improved by using a porous copper electrode of larger surface area.

3. Conclusion

Porous copper electrode has been prepared by copper electrodeposition onto nickel modified graphite felt according to an original process combining an optimized galvanostatic method and flowing conditions. Low concentrations of metal salts and high applied current were necessary to achieve a homogeneous metallic coating of the fibres. These results demonstrate the possibility to deposit a new metal layer on already metalized graphite felt and open the door for the preparation of metallic felt of different nature such as cobalt, tin and bismuth. One of the interests of these porous materials is their catalytic properties. Thus, the use of copper modified graphite felt as a cathode in an electrochemical flow cell allowed a selective and quantitative reduction of nitrate solutions of high concentration ($3\text{g NO}_3\text{L}^{-1}$) into NH_4^+ . A maximum NH_4^+ selectivity ($\%S_{\text{NH}_4^+}$) of 96 % and a high current efficiency for NH_4^+ formation ($\% \text{CE } 72 \%$) were recorded with a solution of pH 7.2, a flow rate of 2 mL min^{-1} and a ratio ($i_{\text{app}}/i_{\text{th}}$) equal to 1. These results are suitable for nitrate treatment on a larger scale application.

In summary in this report we set up a novel electrochemical process allowing an efficient conversion of very concentrated nitrate solution to ammonium in neutral medium.

The feasibility of the use of the ammonium solution, obtained by the present process, to produce biohydrogen in dark fermentation via heat treated anaerobic activated sludge will be carried out in continuity of the present work.

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Figure Captions

Fig. 1. Schematic diagram of the flow cell, a, porous electrode; b, counter electrodes (platinum grids); c₁, cationic membrane; c₂, anionic membrane; d, electrolyte inlet; e, electrolyte outlet ; PW, power supply.

Fig. 2. SEM images of metal coated graphite felts

a - Nickel coated graphite felt

b, c, d - Copper electrodeposited on nickel modified graphite felt

Fig. 3. Effect of current and pH on current efficiency.

■ pH 4.8

▲ pH 7.2

Table 1

Summary of optimal experimental conditions for copper plating on nickel modified graphite felt

Parameters	Optimal conditions
[Cu ²⁺]	3×10^{-2} M
initial pH	4-5
Flow rate	5 mL min ⁻¹
Applied current	0.1 Acm ⁻³
On pulse	15 s
Off pulse	35 s
% CE	50 %

Table 2

Experimental conditions to achieve the electrodeposition of copper on nickel modified graphite felt

Experiment	[Cu ²⁺] mol L ⁻¹	On pulse t _{on}	Off pulse t _{off}	t _{off} /t _{on}	On pulse total time	Full electrolysis Time	Electrode aspect	%CE
1	10 ⁻²	7 s	10 s	1.4	2 h 20 min	5 h 40 min	Poor deposit	24
2	3 × 10 ⁻²	7 s	10 s	1.4	2 h 20 min	5 h 40 min	Metallic crust	65
3	10 ⁻²	10 s	10 s	1	6 h 11 min	13 h	Metallic crust	13
4	3 × 10 ⁻²	15 s	35 s	2.3	4 h	13 h	Homogeneous film of cooper	58

Table 3

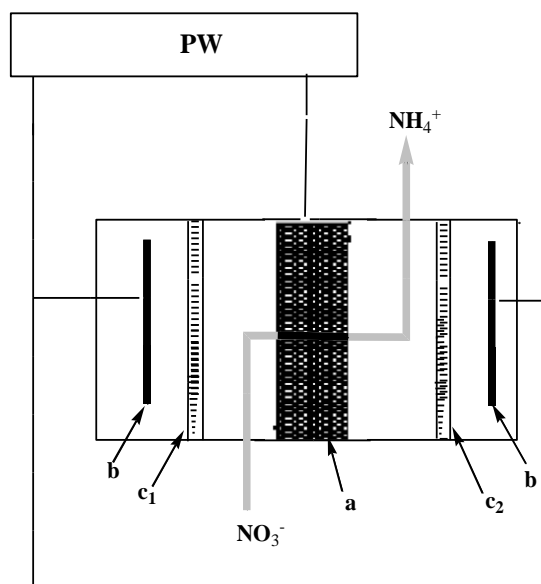
Effect of flow rate and pH on the electrochemical reduction of nitrate to ammonium. (Initial nitrates concentration = $3 \text{ g NO}_3^- \text{ L}^{-1}$; electrolyzed volume = 1liter).

Flow rate (mL min^{-1})	Full electrolysis Time	pH 4.8			pH 7.2		
		% τ	% $\text{S}_{\text{NH}_4^+}$	% $\text{X}_{\text{NO}_2^-}$	% τ	% $\text{S}_{\text{NH}_4^+}$	% $\text{X}_{\text{NO}_2^-}$
1	16 h 40 min	80	71	1	75	86	0.5
2	8h 20 min	70	75	3.5	73	96	1
3	5h 30 min	76	82	2	63	92	3

Table 4

Effect of current and pH on the electrochemical reduction of nitrate to ammonium. ($R_i = i_{app} / i_{th}$; initial nitrate concentration = 3 g $\text{NO}_3^- \text{L}^{-1}$; flow rate = 2 mL min^{-1} ; $i_{th} = 1.24 \text{ A}$).

i_{app} (A)	R_i	pH 4.8				pH 7.2			
		% τ	% τ_{app}	% $\text{S}_{\text{NH}_4^+}$	% $\text{X}_{\text{NO}_2^-}$	% τ	% τ_{app}	% $\text{S}_{\text{NH}_4^+}$	% $\text{X}_{\text{NO}_2^-}$
0.64	0.5	50	100	40	1.5	46	92	60	0.5
0.93	0.75	68	90	60	2.5	66	88	79	0.5
1.24	1	70	70	75	3.5	73	73	96	1
1.55	1.25	75	75	68	3.5	79	79	96	1.5
1.86	1.5	82	82	62	4	83	83	93	1.5

**Fig.1**

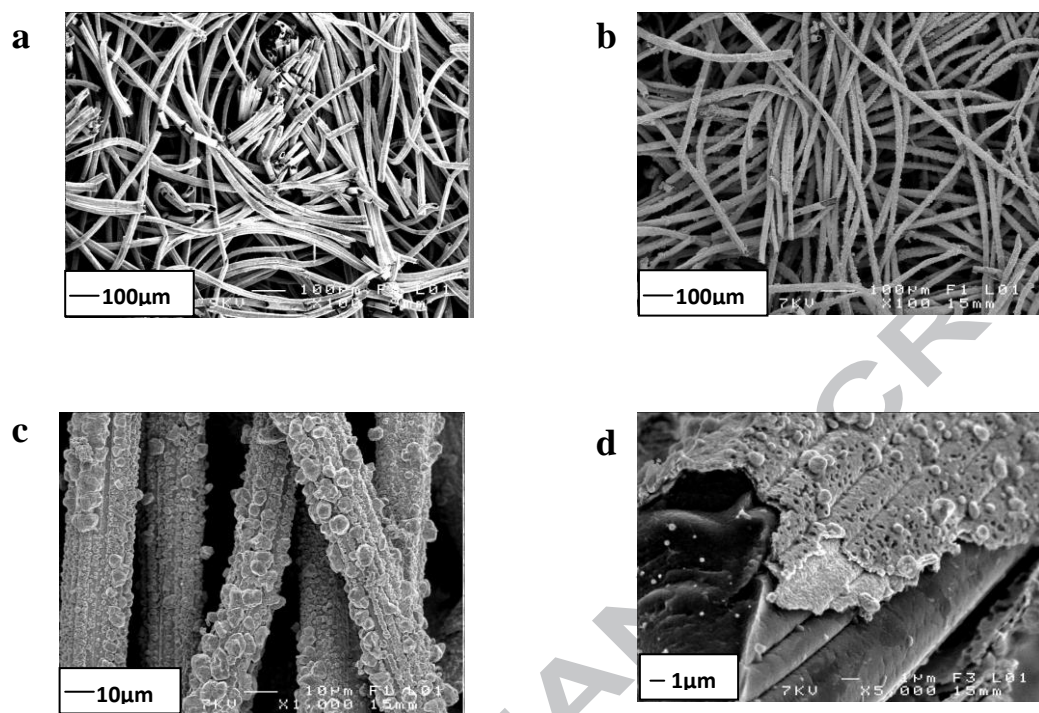


Fig. 2.

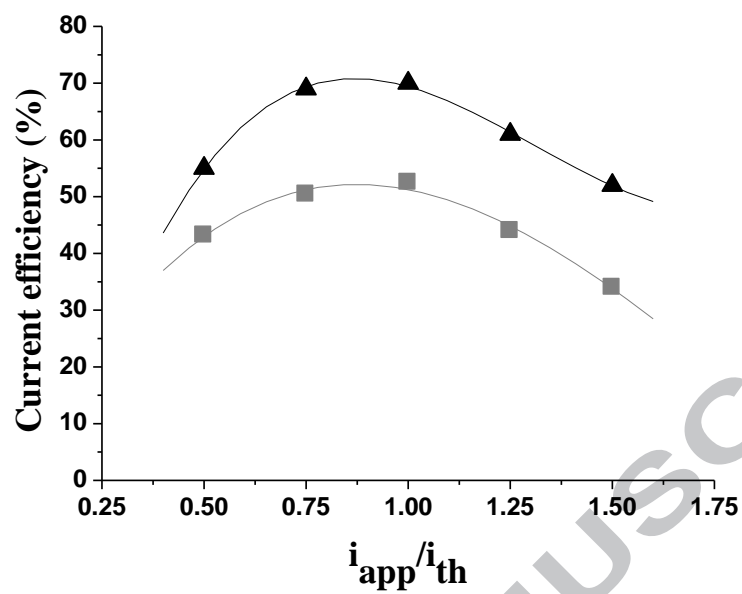


Fig. 3.

Research Highlights

- A novel porous copper electrode of high surface areas and low densities was prepared
- The copper cathode was examined for electroreduction of nitrate to ammonium
- selective and quantitative reduction of nitrate was obtained in neutral solution
- High chemical yields and current efficiencies toward ammonium formation was recorded
- Recycling of nitrates as an abundant and inexpensive source of ammonium